

Wilson – Chemistry Problem-solving – 1st Edition

Errata

Page 26 – Dimensional Analysis – End of second paragraph: double-use of “only” (remove first only)

Pages 96-97 – Formula Mass – Worked Example: there should be 2 H atoms and 1 O atom (*NOT* one H atom and two O atoms)

Page 132 – Balancing Chemical Equations – End of first paragraph: oxygen gas (*NOT* chlorine gas)

Page 137 – Balancing Chemical Equations – Problem 15: typo – should be right-facing arrow

Page 144 – Stoichiometry Calculations – Worked Example 3: 8.395×10^{23} (*NOT* 6.395×10^{23})

Page 145 – Stoichiometry Calculations – Worked Example 3: Step 04: Na⁺ ions (*NOT* H₂O)

Page 185 – Molarity Calculations – Worked Example 3: missing the number 3 after Worked Example

Page 186 – Molarity Calculations – Worked Example 3: error in Step 02 – Solve portion: problem incorrectly uses 1 mol Cl⁻ instead of 2 mol Cl⁻ – this causes the entire problem to be solved incorrectly. Answer for this step should be 0.0054 mol Cl⁻.

Page 187 – Molarity Calculations – Worked Example 3: final answer should be 1.47×10^{-1} M Cl⁻ sol'n

Page 205 – Molarity Calculations – It's Your Turn – Problem 8: it should say “If Yes, what are the conversion factors for this step?” and there should be two lines instead of one

Page 211 – Molarity Calculations – Additional Practice Problem 9: balanced chemical equation should be: $2 \text{AgNO}_3(aq) + \text{CaCl}_2(aq) \rightarrow 2 \text{AgCl}(s) + \text{Ca}(\text{NO}_3)_2(aq)$
(*NOT* $2 \text{AgNO}_3(aq) + \text{CaCl}_2(aq) \rightarrow \text{AgCl}(s) + 3 \text{Ca}(\text{NO}_3)_2(aq)$)

Page 215 – Appendix B – Module 4 – 2: 6.70616629×10^8 mi/hr (*NOT* 6.70343384×10^8 mi/hr)

Page 218 – Appendix C – Module 4 – 6: 4.0074×10^4 km (*NOT* 4.0091×10^4 km)

Page 218 – Appendix C – Module 4 – 7: 7.61207×10^2 mi/hr (*NOT* 7.60901×10^2 mi/hr)

Page 218 – Appendix C – Module 4 – 3: 1.36×10^4 μL (micro unit is missing)

Page 218 – Appendix C – Module 5 – 1e: beryllium hydrogen sulfite (*NOT* sulfate)

Page 218 – Appendix C – Module 5 – 2d: $\text{K}_2\text{Cr}_2\text{O}_7$ (*NOT* K_2CrO_4)

Page 219 – Appendix C – Modules 5 and 6: Module 5's answers end prematurely and are completed in Module 6 (formatting error)

Page 219 – Appendix C – Module 6 – 1b: 127.11 amu (*NOT* 68.06 amu)

Page 219 – Appendix C – Module 6 – 1c: 399.64 amu (*NOT* 396.61 amu)

Page 219 – Appendix C – Module 6 – 2b: 127.11 g/mol (*NOT* 68.06 g/mol)

Page 219 – Appendix C – Module 6 – 2c: 399.64 g/mol (*NOT* 396.61 g/mol)

Page 219 – Appendix C – Module 7 – 12: 3,1,3,1 (*NOT* 1,3,3,1)

Page 219 – Appendix C – Module 8 – 3: 1.65×10^1 g $(\text{NH}_4)_2\text{HPO}_4$ (*NOT* 1.63×10^1 g $(\text{NH}_4)_2\text{HPO}_4$)

Page 219 – Appendix C – Module 8 – 8: 3.82×10^1 g $\text{Al}(\text{OH})_3$ (*NOT* 2.16×10^1 g $\text{Al}(\text{OH})_3$)

Page 219 – Appendix C – Module 8 – 9: 3.765 mol Fe^{3+} (*NOT* 5.647 mol Fe^{3+})

Page 219 – Appendix C – Module 8 – 10: 8.786×10^2 g $\text{Ca}_3(\text{PO}_4)_2$ (*NOT* 2.639×10^3 g $\text{Ca}_3(\text{PO}_4)_2$)

The author offers his sincere apologies for the inconvenience these errors have caused.